

GENERAL CHEMISTRY I (CHE1052) FALL 2019 SYLLABUS

INSTRUCTOR: Dr. Ken Martin
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Office Hours: MWF 10:00 – noon,
R 11:00 – noon,
and by appointment

LECTURE (4 units): Section 1 MWF 8:30 am – 9:35 am Taylor 106

REVIEW SESSIONS: MTWR 5:00 – 6:00 pm Latter 102

LEARNING MATERIALS*:

1. *Textbook:* Tro, Chemistry: A Molecular Approach Plus MasteringChemistry, Pearson, 4th Edition, ISBN-13: 9780134103976 (hardcover text), 9780134162454 (looseleaf text), or 9780134162485 (etext)
2. *Online Homework:* MasteringChemistry www.masteringchemistry.com (bundled with text or purchased separately) Course ID: **CHE1052FALL2019**
3. *Course Website:* Canvas canvas.pointloma.edu **CHE1052-1 FA19 – General Chemistry I**
4. *Scientific Calculator:* Non-graphing, non-programmable calculator required for exams and quizzes.
5. *Clicker:* I-Clicker 2, ISBN-13: 9781429280471
6. *Optional Materials:* Tro, Study Guide for Chemistry: A Molecular Approach, Pearson, 4th Edition 2017, ISBN-13: 9780134066271. Tro, Selected Solutions Manual for Chemistry: A Molecular Approach, Pearson, 4th Edition 2017, ISBN-13: 9780134066288.

* These materials are used for both semesters of General Chemistry. No new materials will be required for General Chemistry II (CHE1053).

COURSE DESCRIPTION: Study of the basic principles of modern chemistry. Emphasis on atomic and molecular structure, thermochemistry, gas laws, chemical bonding, states of matter, and solutions. Prerequisite(s): Satisfactory high school background or CHE 1003 or PSC 1014.
Co-requisite(s): CHE 1052L.

LEARNING OUTCOMES: An understanding of chemistry is a necessary part of an education in the basic and applied sciences, engineering, and medical professions. It also provides insight and increased comprehension regarding current events and proposed policies, not to mention a better understanding of the world in which we live.

Specifically, upon completion of this course, students will be able to:

- Demonstrate a foundational knowledge of the general principles of chemistry including atomic and molecular structure, thermochemistry, chemical bonding, states of matter, and the behavior of solutions.
- Solve problems related to unit conversions, stoichiometry, energy calculations, and gas laws.
- Perform basic chemical laboratory techniques related to the topics listed above.

General Education Learning Outcome 1e Quantitative Reasoning: Students will be able to solve problems that are quantitative in nature. This learning outcome will be assessed directly using problems on the final exam that are quantitative in nature.

EVALUATION: The activities described below will contribute to your overall course grade according to the following:

Hour Examinations (4)	50%
Homework (online & other)	15%
Quizzes and In-Class Activities	15%
Final Examination	20%

Letter grades will be assigned at the end of the course based on your percentage of total possible points, according to the following approximate scale:

A	90 – 100%
B	80 – 89%
C	70 – 79%
D	60 – 69%
NC/F	< 60%

(+) and (–) grades will be assigned within each bracket. (There is no A+ grade.)

Strategies for success in CHE152

1. It is crucial that you primarily understand as well as memorize course material. You will be expected to synthesize your knowledge on homework, quizzes, and exams. Focus on recognizing patterns and learn to apply the problem solving strategies that are introduced in the book and lecture.
2. Working problems is the key to success. Work the practice problems in the book as you read the material and start homework sets early so that you can take advantage of office hours and review sessions.
3. Come prepared to class. The time you invest in reading the assigned sections, taking notes, and answering pre-lecture problems will be repaid in full when it comes time to study for exams.
4. Get help if you do not understand something. I am here for you!

Advice from recent General Chemistry students

- “Stay on top from the beginning. Work on test taking strategy. Study. Exams are everything.”

- “Read the sections assigned for each lecture beforehand. Even if you don't understand what you are reading, it will make so much sense when the instructor explains it. Doing this keeps you on top of this class and makes quizzes and tests much easier to study for.”
- “I advise them to start mastering chemistry assignments as soon as they are assigned. They should take advantage of office hours, too.”
- “Go to office hours if you don't understand something! Read the textbook - it actually helps so much to go over those examples and do them yourself. When studying for the exams, redo examples from class handouts.”
- “Dedicate a certain amount of time each day to reviewing in class material and go to his office hours if you are confused about anything at all.”
- “Pay attention and show up to class! Also read and take notes on the sections that will be covered in lecture. If you're struggling in the class or just not understanding even one thing, go to the tutoring center or office hours and really get it down.”
- “Pay attention in lab, as long as you do well in lab you should be fine.”
- “Study hard and prioritize your time. Also, make a good relationship with the professor.”
- “Apply yourself and work hard. You as an individual determine your success in general chemistry.”
- “Take the class seriously from the very beginning. Review notes directly after class.”

ADMINISTRATION:

1. Attendance: You are responsible for all the material covered during class. Regular and punctual attendance at all classes is considered essential to optimum academic achievement. If the student is absent from more than 10 percent of class meetings, the faculty member can file a written report which may result in de-enrollment. If the absences exceed 20 percent, the student may be de-enrolled without notice until the university drop date or, after that date, receive the appropriate grade for their work and participation. See [Class Attendance](#) in the Undergraduate Academic Catalog.
2. The use of portable electronic devices (phones, laptops, iPods, etc.) not related to the course is not permitted in the classroom.
3. Online Homework: Homework will be assigned regularly through [MasteringChemistry](#) (Course ID: CHE1052FALL2019). Successful completion of the homework is essential in mastering the course material. Late assignments will not be accepted.
4. In-Class Activities: In-Class activities will be assigned and collected periodically during the semester. In-class activities cannot be made up; however, the lowest in-class activity score will be discarded when final grades are computed.
5. Exams and Quizzes: Four exams and a comprehensive final will be given during the semester. Make-up exams will be arranged only if the instructor is contacted prior to the scheduled exam time and then only if you present an institutionally valid excuse. Unannounced quizzes will be given periodically throughout the semester. Quizzes cannot be made up; however, the lowest quiz score will be discarded when final grades are computed. *Only non-graphing and non-programmable calculators may be used for exams and quizzes.*
6. Course Website: [Canvas](#) (CHE1052-1 FA19 – General Chemistry I) is used as a repository for course material such as lecture notes, slides, and miscellaneous items. Announcements will be sent out via Canvas. It is your responsibility to check Canvas regularly and to confirm that your correct email address is in the system. Grades will be posted periodically to Canvas.

OTHER MATTERS (OFFICIAL PLNU POLICIES):

Academic Accommodations Policy: While all students are expected to meet the minimum standards for completion of this course as established by the instructor, students with disabilities may require academic adjustments, modifications or auxiliary aids/services. At Point Loma Nazarene University (PLNU), these students are requested to register with the Disability Resource Center (DRC), located in the Bond Academic Center. (DRC@pointloma.edu or 619-849-2486). The DRC's policies and procedures for assisting such students in the development of an appropriate academic adjustment plan (AP) allows PLNU to comply with Section 504 of the Rehabilitation Act and the Americans with Disabilities Act. Section 504 (a) prohibits discrimination against students with special needs and guarantees all qualified students equal access to and benefits of PLNU programs and activities. After the student files the required documentation, the DRC, in conjunction with the student, will develop an AP to meet that student's specific learning needs. The DRC will thereafter email the student's AP to all faculty who teach courses in which the student is enrolled each semester. The AP must be implemented in all such courses. If students do not wish to avail themselves of some or all of the elements of their AP in a particular course, it is the responsibility of those students to notify their professor in that course. PLNU highly recommends that DRC students speak with their professors during the first two weeks of each semester about the applicability of their AP in that particular course and/or if they do not desire to take advantage of some or all of the elements of their AP in that course.

Academic Honesty Policy: Students should demonstrate academic honesty by doing original work and by giving appropriate credit to the ideas of others. Academic dishonesty is the act of presenting information, ideas, and/or concepts as one's own when in reality they are the results of another person's creativity and effort. A faculty member who believes a situation involving academic dishonesty has been detected may assign a failing grade for that assignment or examination, or, depending on the seriousness of the offense, for the course. Faculty should follow and students may appeal using the procedure in the university Catalog. See [Academic Honesty](#) for definitions of kinds of academic dishonesty and for further policy information.

Final Examination Policy: Successful completion of this class requires taking the final examination **on its scheduled day**. The final examination schedule is posted on the [Undergraduate Records](#) site. No requests for early examinations or alternative days will be approved.

PLNU Copyright Policy: Point Loma Nazarene University, as a non-profit educational institution, is entitled by law to use materials protected by the US Copyright Act for classroom education. Any use of those materials outside the class may violate the law.

PLNU Mission: Point Loma Nazarene University exists to provide higher education in a vital Christian community where minds are engaged and challenged, character is modeled and formed, and service is an expression of faith. Being of Wesleyan heritage, we strive to be a learning community where grace is foundational, truth is pursued, and holiness is a way of life.

General Education Mission: PLNU provides a foundational course of study in the liberal arts informed by the life, death, and resurrection of Jesus Christ. In keeping with the Wesleyan tradition, the curriculum equips students with a broad range of knowledge and skills within and across disciplines to enrich major study, lifelong learning, and vocational service as Christ-like participants in the world's diverse societies and culture.

CHEMISTRY 1052/1052L CLASS AND LAB SCHEDULE

DATE	LECTURE TOPICS	READING ASSIGNMENT	LABORATORY
9/4 (W)	Introduction, Atoms, Molecules, Scientific Approach, Classification of Matter, Physical/Chemical Changes and Properties	1.1 – 1.4	No Lab
9/6 (F)	Energy, Units/Reliability of Measurements and Solving Chemical Problems	1.5 – 1.8	
9/9 (M)	Atoms, Modern Atomic Theory and Laws, Electrons, Atomic Structure and Subatomic Particles	2.1 – 2.6	Check-in AND Scientific Measurements (<i>p. 1</i>)
9/11 (W)	Periodic Law, Periodic Table, Atomic Mass and Molar Mass	2.7 – 2.9	
9/13 (F)	Chemical Bonds, Chemical Formulas, Molecular Models, Elements and Compounds	3.1 – 3.4	
9/16 (M)	Ionic Compounds, Molecular Compounds, Formula and Molar Mass of Compounds	3.5 – 3.8	Determination of a Chemical Formula: The Reaction of Iodine with Zinc (<i>p. 19</i>) AND Nomenclature Worksheet (<i>handout</i>)
9/18 (W)	Composition of Compounds, Determining Chemical Formulas and Writing/Balancing Chemical Equations	3.9 – 3.11	
9/20 (F)	EXAM #1 (day 8)	Chapters 1 – 3	
9/23 (M)	Reaction Stoichiometry, Limiting Reactant, Theoretical and Percent Yield	4.1 – 4.3	A Cycle of Copper Reactions (<i>p. 27</i>)
9/25 (W)	Solution Concentration and Stoichiometry, Aqueous Solutions and Solubility	4.4 – 4.5	
9/27 (F)	Precipitation Reactions and Representing Aqueous Reactions	4.6 – 4.7	
9/30 (M)	Acid–Base and Gas Evolution Reactions	4.8	Volumetric Titration of Acids and Bases (<i>p. 39</i>)
10/2 (W)	Oxidation–Reduction Reactions	4.9	
10/4 (F)	Pressure and Gas Laws	5.1 – 5.5	
10/7 (M)	Gas Mixtures and Gas Stoichiometry	5.6 – 5.7	Determining R Using a Metal-HCl Reaction (<i>p. 47</i>) AND Oxidation-Reduction Worksheet (<i>handout</i>)
10/9 (W)	Kinetic Molecular Theory, Diffusion, Effusion, and Real Gases	5.8 – 5.10	
10/11 (F)	EXAM #2 (day 17)	Chapters 4 – 5	
10/14 (M)	Energy Definitions, Internal Energy, Heat and Work	6.1 – 6.3	The Molar Volume of Dioxygen and Other Gases (<i>p. 55</i>)
10/16 (W)	Heat and Work, Calorimetry	6.4 – 6.5	
10/18 (F)	Enthalpy and Calorimetry	6.6 – 6.7	
10/21 (M)	Enthalpy of Reaction	6.8 – 6.10	Enthalpy Changes in Chemical Reactions: Hess's Law (<i>p. 65</i>)
10/23 (W)	Quantum Mechanics and Nature of Light	7.1 – 7.2	
10/25 (F)	Fall Break – No Class		

DATE	LECTURE TOPICS	READING ASSIGNMENT	LABORATORY
10/28 (M)	Atomic Spectroscopy, Bohr Model and Wave Nature of Matter	7.3 – 7.4	Absorption and Emission Spectroscopy (p. 77)
10/30 (W)	Quantum Mechanics and Atomic Orbitals	7.5 – 7.6	
11/1 (F)	Periodic Table and Electron Configurations	8.1 – 8.4	
11/4 (M)	Quantum Mechanical Model, Periodic Trends and Ions	8.5 – 8.7	Determination of a Chemical Formula by Titration: The Reaction of Calcium with Water (p.85)
11/6 (W)	Electron Affinities, Metallic Character and Periodic Chemical Behavior	8.8 – 8.9	
11/8 (F)	EXAM #3 (day 28)	Chapters 6 – 8	
11/11 (M)	Types of Chemical Bonds, Lewis Structures, Ionic Bonding and Covalent Bonding	9.1 – 9.5	Ionic and Covalent Bonding: Conductivity of Solutions of Ionic and Covalent Compounds (p. 93)
11/13 (W)	Electronegativity, Bond Polarity, Lewis Structures, Resonance and Formal Charge	9.6 – 9.8	
11/15 (F)	Exceptions to Octet Rule, Bond Energies, Bond Lengths and Metallic Bonds	9.9 – 9.11	
11/18 (M)	VSEPR Theory	10.1 – 10.3	Writing Lewis Structures (p. 105) AND Models of Molecular Shapes: VSEPR Theory & Orbital Hybridization (p. 123)
11/20 (W)	Molecular Geometries	10.4	
11/22 (F)	Molecular Shape, Polarity	10.5	
11/25 (M)	Valence Bond Theory	10.6 – 10.7	Thanksgiving Recess No Lab
11/27 (W)	Thanksgiving Recess – No Class		
11/29 (F)	Thanksgiving Recess – No Class		
12/2 (M)	Molecular Orbital Theory	10.8	Crystal Structures Lab (p. 163) <i>Read Ch. 12.1 – 12.6</i>
12/4 (W)	EXAM #4 (day 37)	Chapters 9 – 10	
12/6 (F)	Intermolecular Forces	11.1 – 11.3	
12/9 (M)	Intermolecular Forces, Vaporization, Vapor Pressure	11.4 – 11.5	Liquids and Solids: The Vapor Pressure and Enthalpy of Vaporization of Water, The Enthalpy of Fusion of Water (p. 149) AND Check-Out
12/11 (W)	Sublimation, Fusion and Phase Diagrams	11.6 – 11.9	
12/13 (F)	Final Exam Review		
12/18 (W)	COMPREHENSIVE FINAL EXAM All Sections Wednesday 4:30 – 7:00 pm <i>(See Final Exam Schedule)</i>	Chapters 1 – 12	